## 3 Choice of Theoretical Method

## **Exercise 1 Determine the Proton Affinity for Pyridine using PM3.**

Open GaussView.
Construct pyridine by double clicking the Ring Fragment icon (1 <sup>st</sup> row, 2 <sup>nd</sup>
across), clicking the pyridine fragment ( $2^{nd}$ row, $2^{nd}$ across) and
clicking in the View workspace. $C_5H_5N$ appears.
Click Calculate / Gaussian and choose Job: Opt+Freq and defaults; Method: pyridine
Semiempirical, PM3, and defaults, Title: pyridine PM3; Link 0:
PYRIDINEPM3.chk and defaults; defaults for other tabs.
Click Submit. Save as PYRIDINEPM3.gjf. Click OK.
After the job is completed, click Yes twice and open PYRIDINEPM3.chk. Click Results /
Summary to see the summary of the calculations.
Close everything but the main Control Panel.
Open the PYRIDINEPM3.log file using a text editor such as WordPad and scroll to the
bottom. Scroll up about 300 lines ( $\sim$ 10 clicks on the scroll bar) and locate the
line "Sum of electronic and thermal Enthalpies". Record this value $H(C_5H_5N) =$
Eh.
Close the log file.
Click File / New / Create MolGroup.
Click the Ring Fragment icon and choose benzene (1 <sup>st</sup> row, 1 <sup>st</sup> across) and click in the
View workspace. $C_6H_6$ appears.
Click the Element Fragment icon twice and choose N. Click N atom (1 <sup>st</sup> across) and
click one of the C atoms in the benzene. $C_5H_5NH^+$ appears.
Optimize at the PM3 level as above. Save as HPYRIDINEPM3.chk and
HPYRIDINEPM3.gjf. Be sure that the charge is 1 in the Method window.
Close everything but the main Control Panel.
Open the HPYRIDINEPM3.log file and record the value of "Sum of electronic and
thermal Enthalpies" $H(C_5H_5NH^+) = $ Eh.
Given the "Sum of electronic and thermal Enthalpies" of $H^+$ is $H(H^+) = 0.002368$ Eh
(note: no electronic energy!), calculate the proton affinity $\Delta_{\mu}H = PA$ for
$C_5H_5N + H^+ \rightarrow C_5H_5NH^+$
using PA = { $H(C_5H_5NH^+) - [H(C_5H_5N) + H(H^+)]$ }(625.5095) = kcal
mol <sup>-1</sup> .
The literature value is $-219.2 \pm 1.7$ kcal mol <sup>-1</sup> . Calculate the percent difference =
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## Exercise 2 Determine the Proton Affinity for Pyridine using Hatree-Fock.

Click File / Recent Files / PYRIDINEPM3.chk.

Click Calculate / Gaussian and choose Opt+Freq for the Job; Hartee-Fock and 6-31G(d) for the method and basis set; and give the calculation an appropriate title.

Submit and save the files as PYRIDINEHF.gjf and PYRIDINEHF.chk. The calculations will require a few minutes to complete. Close everything but the main Control Panel. Record the value  $H(C_5H_5N) =$  \_\_\_\_\_ Eh from the .log file. Likewise, open HPYRIDINEPM3.chk and submit the job as HPYRIDINEHF.gjf and HPYRIDINEHF.chk. Close everything but the main Control Panel. Record the value of  $H(C_5H_5NH^+) =$ \_\_\_\_\_Eh. Calculate the proton affinity PA \_\_\_\_\_kcal mol<sup>-1</sup> and the percent difference =

## Exercise 3 Determine the Proton Affinity for Pyridine using B3LYP/6-31G(d).

Click File / Recent Files / PYRIDINEHF.chk.

Click Calculate / Gaussian and choose Opt+Freq for the Job; DFT, B3LYP and 6-31G(d) for the method and basis set; and give the calculation an appropriate title.

Submit and save the files as PYRIDINEDFT.gjf and PYRIDINEDFT.chk.

The calculations will require a few minutes to complete.

Close everything but the main Control Panel.

Record the value  $H(C_5H_5N) =$  \_\_\_\_\_ Eh from the .log file.

Likewise, open HPYRIDINEHF.chk and submit the job as HPYRIDINEDFT.gjf and HPYRIDINEDFT.chk.

Close everything but the main Control Panel.

Record the value of  $H(C_5H_5NH^+) =$ \_\_\_\_\_Eh. Calculate the proton affinity PA \_\_\_\_\_kcal mol<sup>-1</sup> and the percent difference =

Close everything but the main Control Panel.